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= $k_1 [D]_1 X [B]_1 Y$ $k_2 [D]_2 X [B]_2 Y$ Equation 5 In This

Equation K Cancels Out. $[D]_1 = [D]_2$ Because We Are Starting Off With The Same Initial Concentration Of A In Each Trial. The Order Of Species D Which Is X Is Also Constant. Canceling Terms We Have Left: $\text{Rate}_1 \text{Rate}_2 = K[B]_1 Y K[B]_2 Y$ Equation 6 Consider The Following Chemical Reaction: Apr 1th, 2024

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Laws PLAN: We Inspect The Exponents In The Rate Law, Not The Coefficients Of The Balanced Equation, To Find The Individual Orders. We Add The Individual Orders To Get The Overall Reaction Order. (a) The Exponent Of $[NO]$ Is 2 And The Exponent Of $[O_2]$ Is 1, So The Reaction Is Secon Mar 5th, 2024

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2nd Cl 2to FormHCl. 7 Chemistry-I. Figure 1.1: Plot Of Concentration Against T For A Reaction $H_2 + Cl_2 \rightarrow 2HCl$ The Rate Of This Reaction Does Not Depend On Concentration. 1.6 First Order Kinetics A Reaction Of The first Order Is Represented As $X \rightarrow Y$ Where X Is The Reactant And Y The Product. The Rate Of The Reaction Will Be Directly Feb 17th, 2024

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B. Transfer Translational KE To Vibrational Energy To
Weaken/break Reactant Bonds Molecules At A Given
Temperature Possess A KE Distribution: Number Of
Molecules KE $E > E_A$, Activation Energy - Energy
Barrier That Molecules Have To Surmount In Order To
React. Energy Is Needed To Break Reactant Bonds
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