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33.  $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$   
 Actual yield = 1.45g Theoretical yield = 5.95g  
 $\text{KClO}_4 \times 1 \text{ mol } 122.55 \text{ g} \times \frac{2 \text{ mol O}_2}{1 \text{ mol KClO}_4} \times \frac{1 \text{ mol KClO}_4}{32.00 \text{ g}} = 3.11 \text{ g O}_2$   
 %Yield =  $\frac{1.45 \text{ g}}{3.11 \text{ g}} \times 100\% = 46.6\%$

34.  $4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$   
 Percent purity = 62.8%  
 pure  $\text{FeCO}_3$

Apr 8th, 2024

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Chemistry Supplemental Stoichiometry Problems Answer Key

Stoichiometry Is A Chemical Section Involving The Use Of Relationships Between Reagents And / Or Products In A Chemical Reaction To Determine The Desired Quantitative Data. In Greek, Stokhein Means Element And Metricity Means Measurement, Then Stoichiometry Literally Translated Means ...

Mar 4th, 2024

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Chapter 112 Subchapter C Texas Education Agency April 30th, 2018 - Education Rules Concerning Texas Essential

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