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Mass Percentage Of Carbon In Ethanol? Chapter 4 Molar Mass = $2 \times 12.011 + 6 \times 1.008 + 1 \times 15.999 = 46.069 \text{ G/mol}$ % C = $\frac{2 \times 12.011}{46.069} \times 100 \% = 52.14 \%$
Stoichiometry We Can Also Use The Mole Concept To Calculate Mass Relationships

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