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Table C.2 (K) Nitrogen ( N<sub>2</sub>) 100 150 200 250 300 350  
450 550 700 1,000 1,100 1.200 1,300 Oxygen ( Jan  
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Number Of Moles Of All The Components Of The  
Mixture:  $N = N_1 + N_2 + \dots$  It Follows That The Sum Of  
All The Mole Fractions Of The Mixture Must Also Equal  
1.  $X_1 + X_2 + \dots = 1$  The Mass Of Component I In A  
Mixture Is The Product Of The Number Of Moles Of I  
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