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=  $k[D]^1 X[B]^1 Y$   $k[D]^2 X[B]^2 Y$  Equation 5 In This Equation  $k$  Cancels Out.  $[D]^1 = [D]^2$  Because We Are Starting Off With The Same Initial Concentration Of  $A$  In Each Trial. The Order Of Species  $D$  Which Is  $X$  Is Also Constant. Canceling Terms We Have Left:  $Rate_1 = Rate_2 = k[B]^1 Y = k[B]^2 Y$  Equation 6 Consider The Following Chemical Reaction: Mar 1th, 2024

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So, To Account For Effective Collisions, Another Factor, P Called The Probability Or Steric Factor Is Introduced. So, Rate =  $PZ AB E-E A /RT$  The Activated Complex Theory Or Transition State Theory Reactants  $\rightleftharpoons$  Activated Complex  $\rightarrow$  Products This Theory Is Based On The Fact That Jan 1th, 2024

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