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 33. $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$ Actual yield = 1.45g

$\text{Theoretical yield} = 5.95 \text{ g KClO}_4 \times \frac{1 \text{ mol}}{122.55 \text{ g}} \times \frac{2 \text{ O}_2}{1 \text{ KClO}_4} \times 32.00 \text{ g/mol} = 3.11 \text{ g O}_2$
 $\% \text{Yield} = \frac{1.45 \text{ g}}{3.11 \text{ g}} \times 100\% = 46.6\%$
 $34. \text{ } 4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$
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