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33. $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$ Actual yield = 1.45g
Theoretical yield = $5.95\text{g KClO}_4 \times \frac{1\text{mol}}{122.55\text{g}} \times \frac{2\text{mol O}_2}{1\text{mol KClO}_4} \times 32.00\text{g/mol} = 3.11\text{g O}_2$
%Yield = $\frac{1.45\text{g}}{3.11\text{g}} \times 100\% = 46.6\%$
34. $4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$
Percent purity = 62.8%? pure FeCO_3 4th,
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 O_2 Molecules Should Be Supplied To Burn 1 Mol Of CH_4 ...
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