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33. $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$ Actual yield = 1.45g

The yield = $5.95 \text{ g KClO}_4 \times \frac{1 \text{ mol}}{122.55 \text{ g}} \times 2 \text{ O}_2 \times \frac{1 \text{ mol KClO}_4}{4 \text{ mol O}_2} \times 32.00 \text{ g/mol} = 3.11 \text{ g O}_2$
% Yield = $\frac{1.45 \text{ g}}{3.11 \text{ g}} \times 100\% = 46.6\%$
 $34.4 \text{ g FeCO}_3 + \text{O}_2 \rightarrow 2 \text{ Fe}_2\text{O}_3 + 4 \text{ CO}_2$
Percent purity = 62.8% ?
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