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Mole Fraction Sugar Of Solution χ Sugar = $(0.292 \text{ Mol}) / (0.292 + 55.6) = .00522$ Mole Fraction Water Of Solution χ Water = $(55.6 \text{ Mol}) / (0.292 + 55.6) = 0.995$ (or $1 - \chi$ Sugar) Molality $M = (.292 \text{ Mol}) / (1 \text{ Kg}) = 0.292 \text{ Mol Kg}^{-1}$ Molarity $M = (.292 \text{ Mol}) / (1 \text{ L}) = 0.292 \text{ Mol L}^{-1}$
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3.mole Fraction = 0.950 4.mole Fraction = 0.0360 5.mole Fraction = 0.925
Nsci140802exam 018 10.0points 3.3 G Of NaCl (58.4 G/mol) And 9 G Of KBr (119.8

G/mol) Were Dissolved In 40 G Of Water (18.0 G/mol). What Is The Mole Fraction Of KBr In The Solution? Nsci140805 019 10.0points How Many G May 4th, 2022

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1. A Solution Is Prepared By Mixing 100.0 G Of Water, H₂O, And 100.0 G Of Ethanol, C₂H₅OH. Determine The Mole Fractions Of Each Substance. 2. The Molality Of An

Aqueous Solution Of Sugar (C₁₂H₂₂O₁₁) Is 1.62m. Calculate The Mole Fractions Of Sugar And Water. 3. Calculate The Molality Of 25.0 Grams Of Apr 2th, 2022

K Is The Molal Freezing Point Depression And M Molality Is ...

As 1 Atm = 1.013 × 10⁵ Pa, The Osmotic Pressure Is $\Pi = (0.0036 \times 1.013 \times 10^5)$ Pa = 365 Pa Hence $C = \Pi / RT = (365 \text{ Pa}) / ((8.314 \text{ M}^3 \text{ Pa K}^{-1} \text{ Mol}^{-1}) \times (293 \text{ K})) = 0.15 \text{ Mol M}^{-3}$ As 1 M³ = 1000 L, This Corresponds To $C = 1.5 \times 10^{-4} \text{ Mol L}^{-1}$ Mar 3th, 2022

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