

## Pogil Kinetic Molecular Theory Answer Key Free Pdf Books

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CHAPTER 12 GASES AND KINETIC-MOLECULAR THEORY  
3 Boyle's Law: The Volume-Pressure Relationship  $\{V \propto 1/P$  Or  $\{V = K (1/P)$  Or  $PV = K$   $\{P_1V_1 = K_1$  For One Sample Of A Gas.  $\{P_2V_2 = K_2$  For A Second Sample Of A Gas.  $\{k_1 = K_2$  For The Same Sample Of A Gas At The Same T.  $\{Thus$  We Can Write Boyle's Law Mathematically As  $P_1V_1 = P_2V_2$  Mar 20th, 2024  
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