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Comment On Notation: In Electrochemistry, It Is Common To Denote Cell Voltage (potential Difference) With E To Denote "electro-motive Force" Or "emf". This Can Cause Confusion With The Electric field Amplitude, Also De Noted E (in All fields Of Science), Which Is The Local Gradient Of The Potential!  $6 K_{Ox} = K_{1 Red} = \exp Ne(\Delta E)$  15th, 2024

### **On The Nernst Equation In Electrochemistry**

Where  $E^0$  is The So-called Formal Potential. This Includes The Ratio  $G_{Ox} = G_{Red}$  Under The Given Experimental Conditions. In Electrochemical Kinetics, For Example, The Rate Coefficient Of The Reaction Rate Of The Charge Transfer Process Is Related To This Formal Potential. If We Can Keep The Ratio  $G_{Ox} = G_{Red}$  Constant In A Given Series 16th, 2024

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The Standard That Has Been Chosen By Convention Is:  $2 H^+(aq) + 2 E^- \rightarrow H_2(g)$   $E^0 = 0.00$  Volts Here The Notation  $E^0$  Is Called The Standard Electrode Potential And Is The Assigned Potential Of The Standard Hydrogen Electrode When The Soncentration Of  $H^+$  Is 1.0 8th, 2024

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