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162 CHAPTER 6: THERMOCHEMISTRY To Convert The Answer To Joules, We Write: 101.3 J 0.18 L Atm 1L Atm  $= - \cdot \times = \cdot \text{W} - 18 \text{ J} 6.17$  An Expansion Implies An Increase In Volume, Therefore W Must Be = 325 J (see The Defining Equation For Pressure-volume Work.) If The System Absorbs Heat, Q Must Be = 127 J. The Change In Energy (internal Feb 5th, 2024)

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